

USING THE PERIODIC TABLE

- _____ 1. What is the atomic number of Boron (B)?
- _____ 2. How many electrons does an atom of lead (Pb) contain?
- _____ 3. How many protons does an atom of Nickel (Ni) contain?
- _____ 4. How many neutrons does an atom of Tin (Sn) contain?
- _____ 5. What is the atomic mass of Mercury?
- _____ 6. What is the atomic mass of Gold?
- _____ 7. How many electron shells does an atom of Copper contain?
- _____ 8. How many electrons are there in the outer shell of Iodine?
- _____ 9. Determine the number of neutrons in an atom if its atomic number is 80 and its atomic mass is 201.
- _____ 10. Determine the number of neutrons in an atom that has an atomic number of 88 and an atomic mass of 226.
- _____ 11. Determine the number of neutrons in an atom having atomic mass of 32 and atomic number of 16.
- _____ 12. Determine the number of protons in Tin.
- _____ 13. What is the atomic mass of Boron?
- _____ 14. Which element has an atomic mass of 85 and an atomic number of 37? Draw it:

- _____ 15. Draw a Bohr model picture of an atom of Manganese.

- _____ 16. What is the atomic number of Iron?
- _____ 17. Find the atomic mass of Potassium.
- _____ 18. Determine the atomic mass of the element depicted in Figure I below.

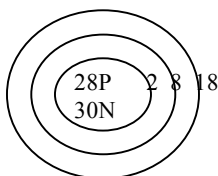


Figure I

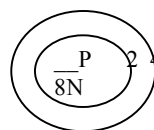


Figure II

- _____ 19. How many protons are in the element depicted in Figure II above?
- _____ 20. Will the element Calcium most likely gain/lose electrons, or stay the same?
- _____ 21. Will the element Sulfur most likely gain or lose electrons & how many?
22. Why are the elements on the right-hand column unlikely to combine with other elements?